

Distribution of Water on Earth (Units of 1015 kg, data from Open University, 1989) Surface water Ground Land: Rivers and Streams Freshwater Lakes 125 Salt Lakes / Inland Seas 104 Ice 79% Glaciers and Ice Caps 29,300 Soil Moisture / Seepage 70 **Ground Water** 8,400 Total On Land 38,000 Salt Water

Overview

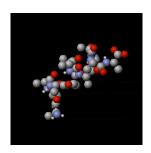
- Special properties of water.
- What makes the oceans move?
 Wind stress coriolis effect thermohaline circulation.
- Interactions between land ocean atmosphere: the carbon cycle.
- Historical perspective: connection between ocean circulation and climate.

Important Properties of Water 1. Polar molecule – high dissolving power Hyaras

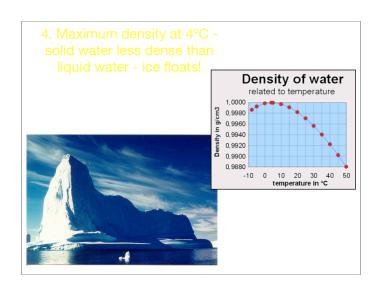
Important dissolved components: NaCl (salt), CO_2



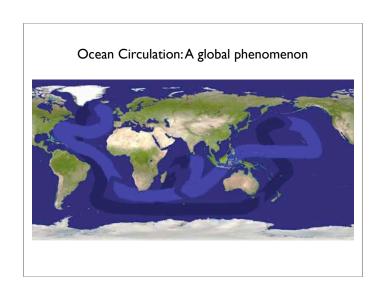
3. High Heat capacity – acts as thermal buffer.



Water has five times the heat capacity of soil: it takes five times the amount of energy to heat up 1 kg of water than it does to heat up 1 kg of soil.







What forces cause ocean water to move?

I.Friction: Results from relative motion between wind and water (wind stress).

2. Coriolis: Results from motion in a rotating coordinate system.

3. Gravity: Gives rise to pressure gradients, buoyancy, and tides.



Prevalent global wind direction is not random!

- determined by the rotation of the Earth.

High

Westerlies

Horse latitudes

By

High

Horse latitudes

By

High

Horse latitudes

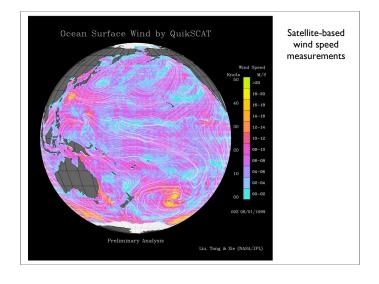
By

High

Horse latitudes

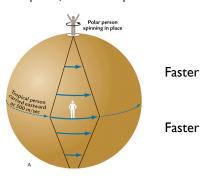
By

High



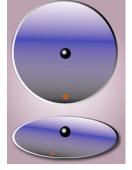
How does the rotation of the Earth affect moving water and air masses ?

→ Different velocity for different latitude: Fastest at the equator, slowest at poles.



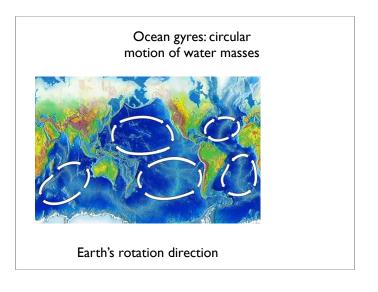
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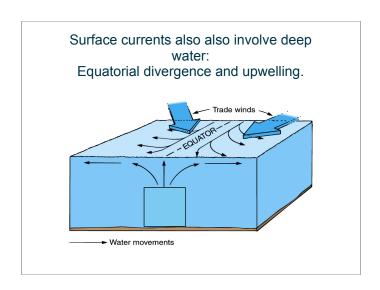
→ Coriolis effect.

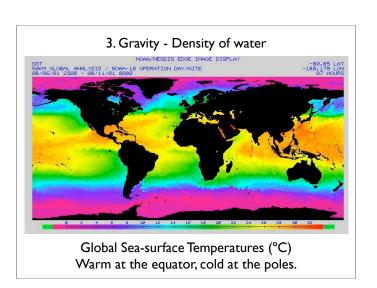


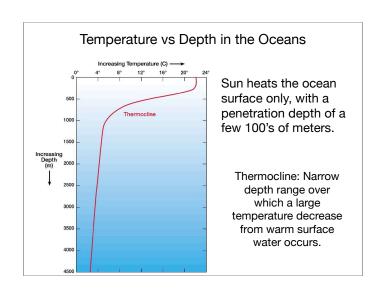
Looking from above : ball moves in a straight line.

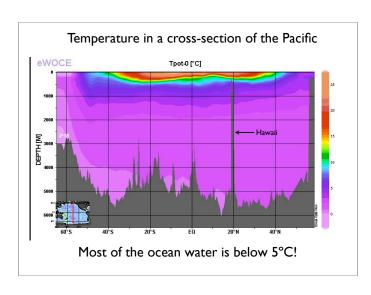
Standing on the rotating plate (fixed - red dot), the ball moves in a curve!

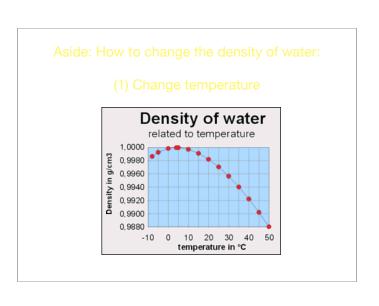






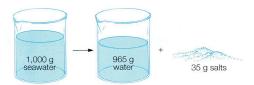






How to change the density of water: (2) Add salt

Density of water increases with increasing amounts of dissolved solids



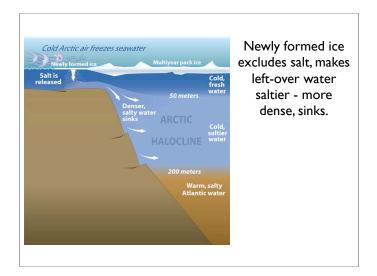
Can also change the relative amount of dissolved material by adding or subtracting water (H₂O): Evaporation, formation of sea ice, river input.

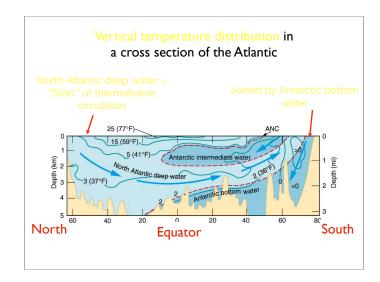
Density driven circulation: Thermohaline circulation

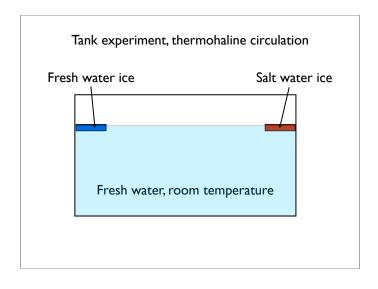
Thermo = temperature

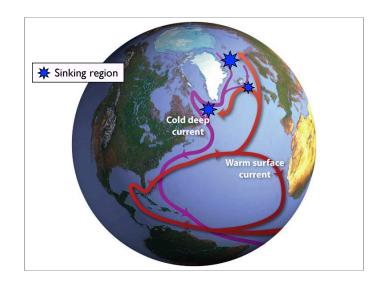
Haline = salt

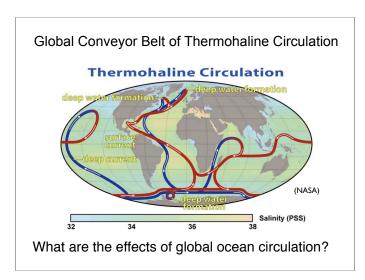
How does the global ocean circulation work?

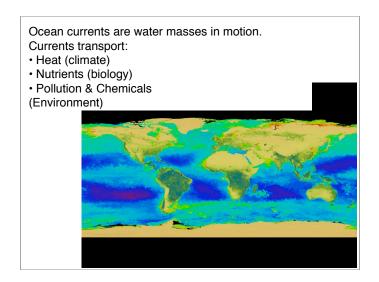












Gulf stream:

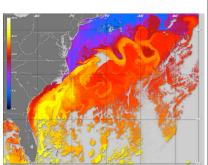
Energy transport: ~ 1.4 petawatts of heat, equivalent to 100 times the world energy demand.

Water transport: up to 80 million m³/s

For comparison: All rivers draining into the Atlantic: ~ 0.6 m³/s

Surface speed: 40 -120 km/day, width 80-150 km, depth: 800-1200 m.

NASA Aqua satellite MODIS image



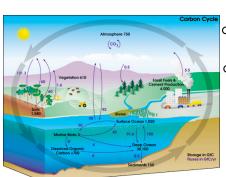
Antarctic Circumpolar Current Water transport: I 35 m³/s. Driven by west winds, unimpeded by land.

Summary so far:

Distribution of water on the Earth
Key properties of water
Three factors driving ocean circulation (Wind,
Earth's rotation, gravity)
Global thermohaline circulation - global transport
of heat (energy) and nutrients

Interaction of ocean, land surface and atmosphere:

The carbon cycle
Carbon chemistry
Weathering reactions
Effect of increasing CO₂ in the atmosphere



Carbon storage in **black** in Gigatons (10⁹ t)

Carbon fluxes (transfer) in purple in Gt/year

Storage: in Atmosphere 750 Gt in Vegetation 610 Gt

Key point: Atmosphere, oceans and solid Earth are part of a linked system!

Carbon chemistry

• Solution:

 $CO_2(atmospheric) \rightleftharpoons CO_2(dissolved)$

• Conversion to carbonic acid: CO_2 (dissolved) + $H_2O \rightleftharpoons H_2CO_3$

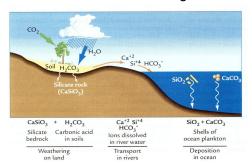
• First ionization:

 $H_2CO_3 \rightleftharpoons H+ + HCO_3^-$ (bicarbonate ion)

• Second ionization:

 $HCO_3^- \rightleftharpoons H+ + CO_3^{2-}$ (carbonate ion)

Chemical weathering



- 1. Weathering reactions involving carbonic acid dissolve silicate minerals.
- 2. Transport of dissolved ions into the ocean.
- 3. Reaction of dissolved $Ca^{2+},\,Mg^{2+}$ to form $CaCO_3$ and $MgCO_3$ sequesters carbon in limestone.

CaCO₃ is used by calcifying organisms, for example coccolithophores, foraminiferans or pteropods

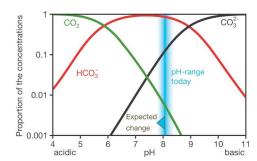






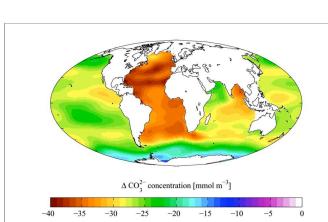
Link between weathering of rocks, CO₂ in the atmosphere and in the ocean: marine life sequesters CO₂!

But: Increased CO₂ content makes oceans more acidic

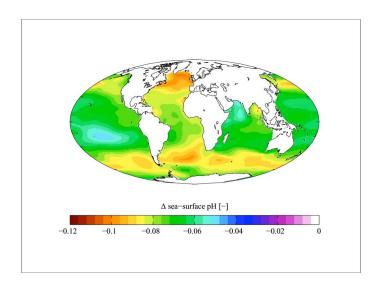


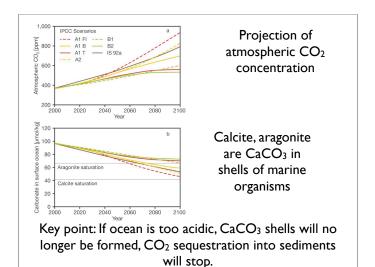
Key point In more acid ocean calcium carbonate shells and skeletons are more difficult to form.

Since these creatures are at the bottom of the food chain, the whole food chain up to baleen whales will be affected.



Change in sea surface carbonate ion (CO_3^{2-}) concentration from the pre-industrial period (1700s) to the present day (1990s).





Historical perspective: What was the climate like in the past?

Is there a connection between climate and ocean circulation?

Past climate inferred from "proxies", e.g. oxygen isotope ratio.

